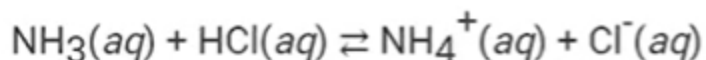


# AP Chemistry Daily Videos

## 4.8 Introduction to Acid-Base Reactions

### Daily Video #1

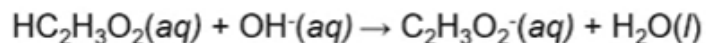
1. How do we define a Bronsted-Lowry acid base reaction?
2. What is a Bronsted-Lowry acid? What is a Bronsted-Lowry base?
3. How can you identify a conjugate acid/base pair?
4. What makes water special (amphiprotic)
5. Pause the video at 3:45 and attempt the problem, then evaluate how you did and identify any errors.



The Brønsted-Lowry bases in the reaction represented above are

- |  |   |
|--|---|
| <input type="radio"/> <b>A</b> $\text{NH}_3(aq)$ and $\text{NH}_4^+(aq)$ | <input type="radio"/> <b>D</b> $\text{HCl}(aq)$ and $\text{NH}_4^+(aq)$ |
| <input type="radio"/> <b>B</b> $\text{NH}_3(aq)$ and $\text{Cl}^-(aq)$   | <input type="radio"/> <b>E</b> $\text{HCl}(aq)$ and $\text{Cl}^-(aq)$   |
| <input type="radio"/> <b>C</b> $\text{NH}_3(aq)$ and $\text{HCl}(aq)$    |   |

6. Pause the video at 4:58 and attempt the problem, then evaluate how you did and identify any errors.



A student performed a titration of  $\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$  with  $\text{KOH}(\text{aq})$ . The net ionic equation for the neutralization reaction that occurs during the titration is shown above.

- a) Identify both of the Bronsted-Lowry conjugate acid-base pairs in the neutralization reaction above. For each pair, label the acid and the base.
- b) In terms of the species in the reaction, explain what occurs during the reaction that identifies it as an acid-base reaction.