

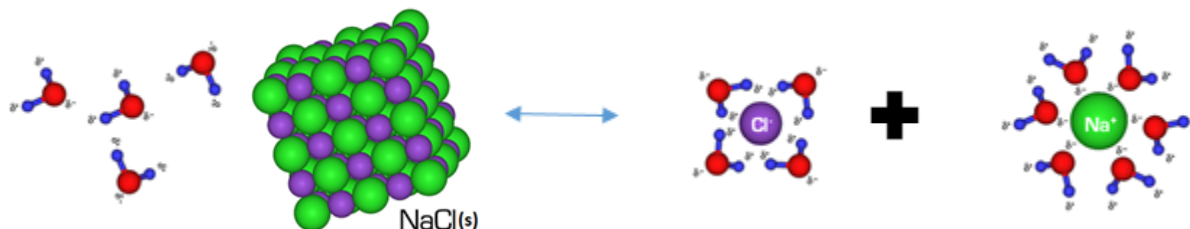
AP Chemistry Daily Videos

7.14 Free Energy of Dissolution

Video #1

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2:40

1. Identify all the bonds/IMF involved when a salt dissolves in water.



Bond/IMF (Identify the type)	Between what particles?	Is the Bond/IMF broken or released?	Is Energy required or released?
	Water Molecules		
	Ions in the salt (Na-Cl)		
	Water and an ion		

Once you evaluate the energy required compared to the energy released you will better understand the enthalpy associated with dissolving this salt.

2. Use pictures to explain what entropy and microstates are.

3. Besides enthalpy what is the other factor needed to determine if a reaction is thermodynamically favorable? Write down the equation that combines these two factors.