

AP Chemistry Daily Videos

7.2 Direction of Reversible Reactions

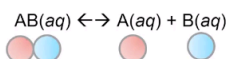
Video #1

1. Draw a picture to represent the following:
 Reactants or the reverse reaction are favored

Products of the forward reaction are favored

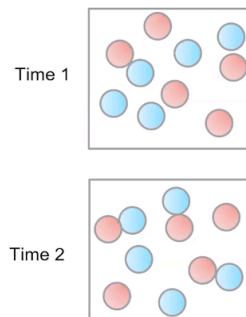


2. Try following the problem before the explanation and answer is given. If you got it wrong, what was your misunderstanding?



A reversible reaction is represented by the equation above. The amount of reactants and products at time 1 and time 2 are shown. Based on the diagram, what can we infer about the rates of the forward and reverse reactions over time?

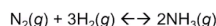
- A) Over time, the forward reaction is favored, because the concentration of AB increases.
- B) Over time, the reverse reaction is favored, because the concentration of A decreases.
- C) Over time, the system reaches equilibrium, because the total number of A and B atoms does not change.
- D) Over time, the system reaches equilibrium, because the reaction stops.



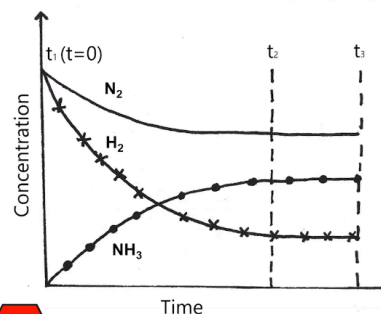
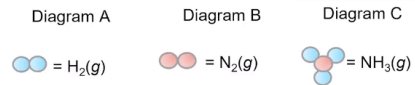
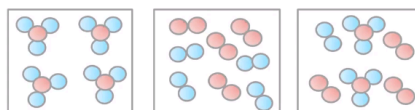
Source: Joanna Scimeca



3. Try following the problem before the explanation and answer is given. If you got it wrong, what was your misunderstanding?



The reaction represented above is allowed to go to completion. What sequence of three of the particle diagrams shown below best represents the progress of the reaction at times t_1 , t_2 , and t_3 ?



Graph: Topic Question 7.1; taken from: AP Classroom

Pause the video and try this problem on your own first.