AP Chemistry Daily Videos

7.5 Magnitude of the Equilibrium Constant

Video #1

1.

Visual	Question	Answer
1/6 3/6	What happens to the value of a fraction as the numerator gets larger?	
1/3	What happens to the value of a fraction as the denominator gets larger?	
	Remember that calculating K, products are in the numerator and reactants are in the denominator. What do you know if K is larger than 1? Draw a picture to the left to represent what happens when K is large.	Which side of the equation is favored?
	What do you know if K is less than one? Draw a picture to the left to represent what happens when K is small.	Which side of the equation is favored?
	What do you know if K=1? Draw a picture to the left to represent what happens when K=1.	Which side of the equation is favored?

2. Ksp is a specific equilibrium constant called the solubility product constant. What does it represent?

3.	Label each image as either: Large Ksp or Small Ksp	
4.	Evaluate how you did in example	1 and identify any errors.
5.	Try to complete this problem	$Ba^{2+}(aq) + EDTA^{4-}(aq) \rightleftharpoons Ba(EDTA)^{2-}(aq)$ K = 7.7 × 10 ⁷
	before the answer is given. Evaluate how you did and identify any errors you made.	The polyatomic ion $C_{10}H_{12}N_2O_8^{4-}$ is commonly abbreviated as EDTA ⁴⁻ . The ion can form complexes with metal ions in aqueous solutions. A complex of EDTA ⁴⁻ with Ba ²⁺ ions forms according to the equation above. A 50.0 mL volume of a solution that has an EDTA ⁴⁻ (aq) concentration of 0.30 M is mixed with 50.0 mL of 0.20 M Ba(NO ₃) ₂ to produce 100.0 mL of solution.
		Considering the value of K for the reaction, determine the concentration of Ba(EDTA) ²⁻ (aq) in the 100.0 mL of solution. Justify your answer.
6.	. What does it mean when an equilibrium reaction is said to go "essentially to completion"?	
7.	Summarize the four key takeawa	ays.
	b.	
	c.	
	d.	

(II) 4:15 (II) 5:26