AP Chemistry Daily Videos

7.8 Representations of Equilibrium

Video #1

1. **What assumptions should you make about particle diagrams?**

2. **How should you approach these diagrams?**

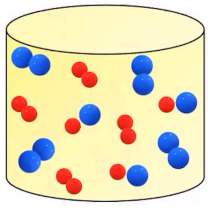
3. **Evaluate how you did and identify any errors you may have made.**

The diagram shows the following reaction at equilibrium in a two-liter reaction vessel. Each particle represents 0.020 moles.

$$Q_2 + Z_2 \rightleftharpoons 2QZ$$

(a) Determine the molarity of Q_2, Z_2, and QZ at equilibrium.

(b) Write the equilibrium constant expression in terms of molarity and calculate the equilibrium constant



2 Liters

Source: Dena K. Leggett