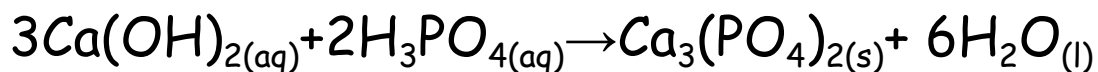


Formal Assessment: Chemical reaction

Name: _____ Period: _____ Date: _____



Consider the chemical reaction mentioned above to answer questions 1-10.

Questions	Answers
1. What are the reactants?	
2. What are the products?	
3. What does (aq) mean in the context of this chemical reaction?	
4. If there is a precipitate in the reaction, please mention it here. If there is no precipitate, just write, there is no precipitate.	
5. What does the (l) mean in the context of this chemical reaction?	
6. What is the total number of oxygen atoms in the product side of this chemical reaction?	
7. How many hydroxides are in $3\text{Ca}(\text{OH})_2$?	
8. How many oxygen atoms are in $3\text{Ca}(\text{OH})_2$?	
9. How many phosphorus atoms are in $\text{Ca}_3(\text{PO}_4)_2$?	
10. How many phosphates are in $\text{Ca}_3(\text{PO}_4)_2$?	

11. What is the molar mass of $\text{Ca}_3(\text{PO}_4)_2$? Make sure you mention units.

Show your work here:

Answer:

12. Balance the following chemical equations and determine which type of reaction they are.

Chemical reaction	Type of chemical reaction
a) $\underline{\hspace{1cm}} \text{H}_2 + \underline{\hspace{1cm}} \text{NO} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} + \underline{\hspace{1cm}} \text{N}_2$	
b) $\underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{CaCl}_2 \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2$	
c) $\underline{\hspace{1cm}} \text{P}_4\text{O}_2 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{P}_2\text{O}_5$	
d) $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$	