

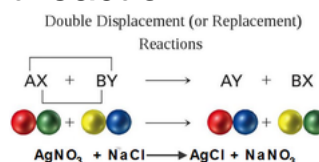


1. **precipitate** A solid that forms from a solution during a chemical reaction.



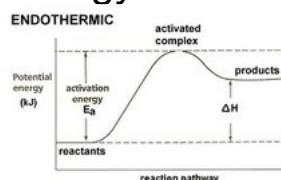
2. **double replacement reaction** a chemical change that involves an exchange of positive ions between two compounds

3. **metathesis reaction** double displacement reaction

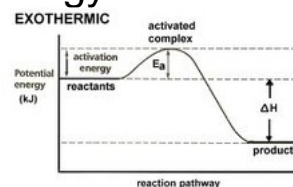


4. **combustion reaction** a chemical reaction that occurs when a substance reacts with oxygen, releasing energy in the form of heat and light

5. **endothermic change** the system absorbs energy from its surroundings



6. **exothermic change** a change in which energy is released

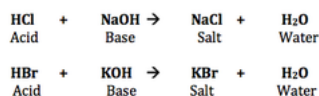


7. **synthesis reaction** a reaction in which two or more substances combine to form a new compound

8. **decomposition reaction** a reaction in which a single compound breaks down to form two or more simpler substances

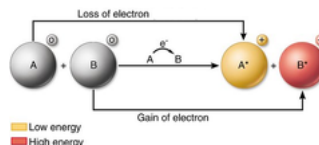


9. **acid-base reaction** a reaction where an acid reacts with a base to produce water and a salt



10. **oxidation number** Positive or negative number that indicates how many electrons an atom has gained, lost, or shared to become stable

11. **oxidation-reduction reaction** a reaction that involves the transfer of electrons between reactants

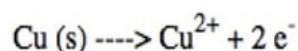


12. **Oxidation** loss of electrons

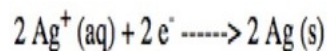
13. **reduction** gain of electrons

14. **single replacement reaction** a chemical change in which one element replaces a second element in a compound

15. **oxidation half reaction** That half of a redox reaction where loss of electrons takes place. In this half, the oxidation number of the reactant atoms increases.



16. **reduction half reaction** the "half" of an oxidation-reduction reaction involving reduction; the half-reaction in which electrons appear as reactants; balanced when each atom type, as well as the charge, is balanced



17. **half-reaction** an equation showing either the oxidation or the reduction that takes place in a redox reaction

18. **neutralization reaction** a reaction in which an acid and a base react in an aqueous solution to produce a salt and water

19. **Molarity** the number of moles of solute per liter of solution



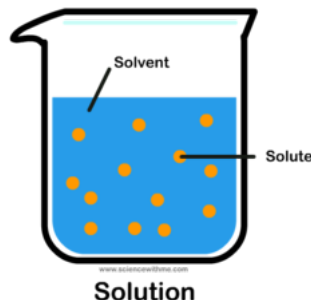
$$\text{Molarity} = \frac{\text{moles of solute}}{\text{volume of solution (L)}}$$

## 20. **Solution Stoichiometry**

A method of calculating the concentration of substances in a chemical reaction by measuring the volumes of solutions that react completely; sometimes called volumetric stoichiometry.

## 21. **Solute**

the substance that is dissolved



## 22. **Solvent**

the substance in which the solute dissolves

## 23. **Solution**

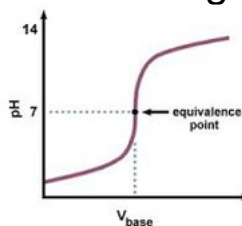
A mixture that forms when one substance dissolves another.

## 24. **Titration**

process in which a solution of known concentration is used to determine the concentration of another solution

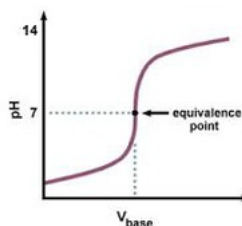
## 25. **titration curve**

graph showing how the pH of a solution changes as acidic and basic solutions are added together



## 26. **equivalence point**

occurs when the moles of acid equal the moles of base in a solution





## AP Chemistry-Chemical Reactions

Study online at [https://quizlet.com/\\_72p44k](https://quizlet.com/_72p44k)

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27. **hydrogen ion (H<sup>+</sup>)** a positively charged ion (H<sup>+</sup>) formed of a hydrogen atom that has lost its electron

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28. **hydronium ion** hydrogen ion combines with a water molecule to form a hydronium ion, H<sub>3</sub>O<sup>(+)</sup>

