

Ionic Compounds with Polyatomic Ions Practice Part 1

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

	Ion Pair	Formula	Name of the compound
1.	$K^+$ $OH^-$		
2	$Ba^{2+}$ $SO_3^{2-}$		
3	$Li^+$ $ClO_3^-$		
4	$Ru^{3+}$ $PO_4^{3-}$		
5	$Ni^{2+}$ $CrO_4^{2-}$		
6	$Cu^+$ $CO_3^{2-}$		
7	$Au^+$ $SO_4^{2-}$		
8	$Au^+$ $PO_3^{3-}$		
9	$Cs^+$ $Cr_2O_7^{2-}$		
10	$Pt^{2+}$ $PO_3^{3-}$		
11	$Al^{3+}$ $CN^-$		

	Ion Pair	Formula	Name of the compound
12	$W^{6+}$ $NO_3^-$		
13	$Pd^{4+}$ $ClO_4^-$		
14	$Ir^{3+}$ $BrO^-$		
15	$Ru^{3+}$ $CO_3^{2-}$		
16	$Ta^{5+}$ $SO_4^{2-}$		
17	$Os^{3+}$ $Cr_2O_7^{2-}$		
18	$Pt^{4+}$ $PO_4^{3-}$		
19	$Pt^{4+}$ $SO_4^{2-}$		
20	$Re^{6+}$ $CrO_4^{2-}$		
21	$Re^{6+}$ $PO_4^{3-}$		
22	$Mo^{6+}$ $SO_3^{2-}$		
23	$Mo^{6+}$ $PO_3^{3-}$		