

Ionic compounds without polyatomic ions practice

Name: _____ Period: _____ Date: _____

	Ion Pair	Formula	Name of the compound
1.	$Be^{2+} S^{2-}$		
2	$Ca^{2+} O^{2-}$		
3	$Al^{3+} F^{-}$		
4	$Ga^{3+} O^{2-}$		
5	$V^{3+} O^{2-}$		
6	$V^{2+} O^{2-}$		
7	$Nb^{+5} S^{2-}$		
8	$Pd^{2+} O^{2-}$		
9	$Au^{+} F^{-}$		
10	$Au^{3+} O^{2-}$		
11	$Pt^{4+} O^{2-}$		

	Ion Pair	Formula	Name of the compound
12	$W^{6+} O^{2-}$		
13	$Os^{4+} O^{2-}$		
14	$W^{6+} N^{3-}$		
15	$Ru^{3+} N^{3-}$		
16	$Re^{7+} O^{2-}$		
17	$Re^{6+} O^{2-}$		
18	$Re^{4+} O^{2-}$		
19	$Cd^{2+} O^{2-}$		
20	$Ag^{+} N^{3-}$		