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1. **Faraday constant** amount of charge per one mole of electrons.

 2. **96,485 C/mol e⁻** Faraday's constant

 3. **Current (ampere)** quantity of charge moving past a point in a circuit per second

 4. **galvanic cell** uses spontaneous chemical reaction to generate electricity

 5. **electrolysis** the process in which a chemical reaction is forced to occur at an electrode by an imposed voltage

 6. **oxidation** loss of electrons (an increase in oxidation number)

 7. **reduction** gain of electrons (a decrease in oxidation number)

 8. **reducing agent** the substance containing the element that gets oxidized

 9. **oxidizing agent** the substance containing the element that gets reduced

 10. **electrochemistry** the study of the interchange of chemical and electrical energy

 11. **galvanic cell** a device in which chemical energy is changed to electrical energy (usually two half-cells) -- the cell potential is positive

 12. **salt bridge** a U-tube filled with an electrolyte or a porous disk in a tube connecting the two half-cells -- allows ions to flow between the two compartments. Keeps cell from having voltage drop to 0

 13. **cathode** the electrode compartment in which reduction occurs (RED CAT)

 14. **anode** the electrode compartment in which oxidation occurs (AN OX)

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standard reduction potentials	the electropotential values corresponding to reduction half-reactions will all solutes at 1 M and all gases at 1 atms
16. battery	a galvanic cell or a group of galvanic cells connected in series -- source of direct current and provide portable power
17. electrolytic cell	uses electrical energy to produce chemical change (nonspontaneously)
18. electrolysis	involves forcing a current through a cell to produce a chemical change for which the cell potential is negative
19. ampere (amp)	unit of current = 1 coulomb of charge per second
20. electrochemical process	any conversion between chemical energy and electric energy
21. electrochemical cell	any device that converts chemical energy into electrical energy or electric energy into chemical energy; consists of redox reactions
22. half-cell	one type of voltaic cell in which either oxidation or reduction occurs
23. salt bridge	a tube containing a strong electrolyte , often potassium sulfate; contain agar; ; allows ions to pass from one half-cell to the other but prevents the solutions from mixing completely; half cells are connected by these
24. electrode	a conductor in a circuit that carries electrons to or from a substance other than a metal
25. battery	a group of voltaic cells connected together
26. Mass increases	at the cathode because aqueous turn into solid
27. Mass decreases	at the anode because solid atoms become aqueous ions



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| 28. oxidation | loss of electrons |
| 29. reduction | gain of electrons |
| 30. electrolyte | a solution that contains ions and can carry a charge |
| 31. redox reaction | name for an oxidation/reduction reaction |
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